

INITIAL CONCENTRATIONS (PART 2, SOLUTION 1)

5.0 ml 0.00200 M Fe^{3+}

1.0 ml 0.00200 M SCN^-

4.0 ml 0.050 M HNO_3

$$[\text{Fe}^{3+}] = \left(\frac{5.0 \text{ ml}}{10.0 \text{ ml}} \right) (0.00200) = 0.00100$$

$$[\text{SCN}^-] = \left(\frac{1.0 \text{ ml}}{10.0 \text{ ml}} \right) (0.00200) = 0.00020$$