Name_____Section_____

CHM115 Lab 3 Report Form
Titration 1- Standardization of NaOH solution

Write the balanced equation for this titration. Circle the analyte.

As always, attach one sample of each type of calculation. Watch sig figs!

	Trial 1	**Trial 2**	**Trial 3**	**Trial 4**	**Trial 5**
Mass pure KHP (g)					
Moles pure KHP					
Volume of NaOH used (mL)					
M NaOH					

Average M _____

	Trial 1	**Trial 2**	**Trial 3**	**Trial 4**	**Trial 5**
Deviation					

Average deviation_____

Be sure to indicate if any data were omitted from the calculations and your reasoning in leaving them out.

Titration 2: Determining the % KHP in an impure sample

Write the balanced equation for this titration. Circle what the analyte is this time.

M NaOH used (from Lab 4) _____

Show one sample of each type of calculation on the back. Watch sig figs!

	Trial 1	**Trial 2**	**Trial 3**	**Trial 4**	**Trial 5**
Volume of NaOH used (mL))					
Moles NaOH					
Moles pure KHP					
Mass pure KHP (g)					
Mass impure KHP used (g)					
% KHP					

Average % KHP _____ Unknown Number_____

	Trial 1	**Trial 2**	**Trial 3**	**Trial 4**	**Trial 5**
Deviation					

Average deviation_____

Be sure to indicate if any data were omitted from the calculations and your reasoning in leaving them out.

Parts per thousand (ppt) is another way of analyzing data. To find ppt, divide the average deviation by average value (% KHP) and multiply by 1000: _____ ppt

A good value for this lab is 3 ppt. Do you think your results are good?

Is this a measure of accuracy or precision? Explain.