

Name \_\_\_\_\_ Section \_\_\_\_\_

**CHM115 Lab 4 Report Form**  
**Determining the % KHP in an impure sample**

Write the balanced equation for this titration. Circle what the analyte is this time.

M NaOH used (from Lab 3) \_\_\_\_\_

Show one sample of each type of calculation on the back. Watch sig figs!

	<b>Trial 1</b>	<b>Trial 2</b>	<b>Trial 3</b>	<b>Trial 4</b>	<b>Trial 5</b>
Volume of NaOH used (mL)					
Moles NaOH					
Moles pure KHP					
Mass pure KHP (g)					
Mass impure KHP used (g)					
% KHP					

Average % KHP \_\_\_\_\_

Unknown Number \_\_\_\_\_

	<b>Trial 1</b>	<b>Trial 2</b>	<b>Trial 3</b>	<b>Trial 4</b>	<b>Trial 5</b>
Deviation					

Average deviation \_\_\_\_\_

Be sure to indicate if any data were omitted from the calculations and your reasoning in leaving them out.

Which THREE trials from the five you completed give the best precision?

Using the THREE trials that give the best precision, compute the average and average deviation below.

%KHP in Unknown: \_\_\_\_\_  $\pm$  \_\_\_\_\_