Name_____Section_____
Partner_____

CHM 111 Purity of KHP Lab Report Form

Write the balanced equation for this titration. Circle the analyte.

M NaOH used (from Lab 3) _____

Show one sample of each type of calculation. Watch sig figs!

	Trial 1	Trial 2	Trial 3	Trial 4	Trial 5
Volume of NaOH used (mL))					
Moles NaOH					
Moles pure KHP					
Mass pure KHP (g)					
Mass impure KHP used (g)					
% KHP					

Average % KHP _____
Unknown Number _____

	Trial 1	Trial 2	Trial 3	Trial 4	Trial 5
Deviation					

Average deviation _____

Be sure to indicate if any data were omitted from the calculations and your reasoning in leaving them out.

Parts per thousand (ppt) is another way of analyzing data. To find ppt, divide the average deviation by average value (% KHP) and multiply by 1000: _____ ppt

A good value for this lab is 3 ppt. Do you think your results are good?

Is this a measure of accuracy or precision? Explain.