

16. Complete the following chart. Items on the left have to be correct to get credit for items on the right. If multiple resonance structures exist, only draw one. Include ALL valence electrons in your diagrams.

Molecule	Lewis Diagram (2 pts)	Molecular Shape ¹ (2 pts)	Polar (Y/N) (1 pt)
XeO ₃			
NO ₂ ⁺			
ClF ₃			
SO ₄ ²⁻			
IF ₅			

¹One of: linear, bent, tetrahedral, trigonal planar, trigonal pyramid, trigonal bi-pyramid, see-saw, T-shaped, octahedral, square pyramid, or square planar.